Q: Define an element	Q: Define a compound.
Q: Define a mixture	Q: Describe the test for oxygen
Q: Describe the test for Hydrogen.	Q: Describe the test for Chlorine gas?
Q: Describe the test for Carbon Dioxide.	Q: Define a Solvent

A: 2 or more different elements chemically bonded together.	A: Made from 1 type of atom
A: Relights a glowing splint.	A: Different elements of compounds that are NOT bonded together.
A: Turns blue litmus paper white (Bleaches it).	A: Burns with a squeaky pop.
A: The liquid that does the dissolving e.g. water.	A: Turns limewater from clear to cloudy.

Q: Define a solvent	Q: Describe a solution
Q: In chromatography, why do you draw the starting line in pencil?	Q: Why is it important the the water level is below the starting line in Chromatography.
Q: In Chromatography, what is the stationary phase?	Q: In Chromatography, what is the mobile phase?
Q: How do you calculate Rf values in Chromatography?	Q: In Chromatography, why do we calculate Rf values?

A: When a solute is dissolved in a solvent.	A: The solid that is dissolved in a solvent. E.g. salt.
A: The ink spots would dissolve in the water.	A: Because ink would dissolve in the water and contaminate the samples.
A: The phase that moves. E.g. the water	A: The phase that does not move e.g. the paper
A: To standardize the result so anyone will get the same result.	A: Distance spot moves / distance water moves

Q: Name the flame test colours for lithium and sodium	Q: Name the flame test colours for potassium, calcium and copper
Q: When testing for Halides (Cl, Br & I), you add silver nitrate and dilute nitric acid. What colour does each Halide turn?	Q: How do you test for a sulfate ion?
Q: How do you test for a carbonate ion?	Q: You can use Sodium hydroxide solution to test for many metal ions. What would you observe for Calcium and Copper?
Q: You can use Sodium hydroxide solution to test for many metal ions. What would you observe for iron (II) and Iron (III) ions?	Q: You can use Sodium hydroxide solution to test for many metal ions. What would you observe for Aluminium and Magnesium ions

A: Lilac –potassium Brick red – Calcium Green - Copper.	A: Red for Lithium and orange/yellow for Sodium.
A: Add Barium Chloride & dilute hydrochloric acid and observe a white percipitate	A: Chloride – White Bromide – Cream Iodide - Yellow
A: Calcium – white precipitate that doesn't dissolve Copper – Blue Precipitate	A: Add dilute acid and observe fizzing. Confirm the gas is CO_2 by testing with limewater
A: Aluminium – White precipitate that dissolves Magnesium – White precipitate that doesn't dissolve	A: Iron (II) – Green precipitate Iron (III) – Brown precipitate